

## Slow Pyrolysis of Sugarcane Bagasse for the Production of Char and the Potential of Its By-Product for Wood Protection

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**Abstract:** Sugarcane bagasse was pyrolyzed using a laboratory fixed bed reactor to produce char and its by-product (pyrolysis liquid). The pyrolysis experiments were carried out using different temperatures (400°C and 500°C), heating rate (1 °C/min and 10 °C/min), and holding time (30 min and 60 min). Char was characterized according to its thermal properties, while the pyrolysis liquid was tested for its anti-fungal and anti-termite activities. Pyrolysis temperature and heating rate had a significant influence on the char properties and the yield of char and pyrolysis liquid, where a high-quality char and high yield of pyrolysis liquid can be obtained at a temperature of 500°C and a heating rate of 10 °C/min. The yield of char and pyrolysis liquid was 28.97% and 55.46%, respectively. The principal compounds of pyrolysis liquid were water, acetic acid, glycolaldehyde, 1-hydroxy-2-propanone, methanol, formic acid, levoglucosan, furfural, followed by some phenol compounds and guaiacol derivatives. Pyrolysis liquid at a concentration of 0.20% and 0.25% (v/v) caused a 100% inhibition of *Coniophora puteana* and *Trametes versicolor*, respectively, when performing inhibition growth tests in Petri dishes. Filter paper treated with 10% of pyrolysis liquid caused 100% of termite mortality, while only 5.65%–7.03% of the treated filter papers consumed by termites at such concentration. Pyrolysis liquid is potentially effective to be used in the formulation of wood protection against fungi and termites.

**Keywords:** Anti-fungal; anti-termite; biomass; char; pyrolysis liquid; slow pyrolysis; sugarcane bagasse; valorization

### 1 Introduction

Valorization, a concept of converting the biomass or waste into more useful products, has great potential in the context of waste management, sustainable practices, and promoting the circular bio-economy system. Among other renewable resources, lignocellulosic biomass offers unique advantages, as an abundant, low-cost resource, especially for the production of fuel, energy, and source of the chemical feedstock. Presently, biomass is the only renewable material that can produce liquid fuels as an alternative to fossil fuel [1]. In



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many developing countries, biomass materials will contribute significantly to the development of sustainable energy sources. However, most often, direct combustion is applied when biomass is used for energy. The said method only generates low efficiency (5%–15%) [2], due to lignocellulosic biomass's natural properties that have low density and low calorific value. The conversion of the biomass residue into the concentrated form of energy can enhance biomass utilization efficiency.

Sugarcane is one of the world's most important crops, representing 21.1% of the total global crop production [3]. During its processing in the mill, sugarcane's by-product, which is called bagasse, still possesses a high potential for bioenergy production. Nevertheless, being a seasonal plant, bagasse needs to be stored if it is to be utilized outside of its peak harvesting time. Also, being a hygroscopic material, bagasse is prone to moisture change during storage and subject to biodegradation [4,5]. To overcome these problems, the thermochemical conversion of bagasse can be applied by transforming it into carbon-rich char, which can subsequently be stored without any energy depletion.

Pyrolysis is a thermal decomposition of organic materials in an oxygen-deficient environment, converting the materials into solid char, condensable liquid, and non-condensable gas [6]. The relative quantity of these three fractions can be controlled by adjusting the pyrolysis parameters, such as temperature, heating rate, and solid residence time [7]. Slow pyrolysis, which occurs at a low heating rate and long vapor residence time, targets char as its main product [6,8]. The usual operational temperatures are between 400°C and 600°C, and typically yields 25–35 wt% of char and 30–50 wt% of liquid product [9]. This technique is attractive in converting the unstable biomass, such as sugarcane bagasse, which contains higher volatile matter to a stable energy-rich product. The char can also be used for electricity production or be upgraded as a high-value product such as activated carbon and soil amendments [9]. The potential of biochar for the soil application is interesting, especially when the feedstock is derived from agricultural residues as they contain minerals (ash) withdrawn from the soil. Soil amendment was reported results in a carbon-negative economy due to the retention of carbon-rich material, which is very stable once stored in soil [10,11].

The additional liquid product can also be recovered, with an interesting balance of the mass yield, yet at a modest energy yield [12]. There are various terms for the liquid product presented in the literature, including pyrolysis liquid [13], pyrolysis distillate [12,14], bio-oil [15], tar oil [16], wood vinegar [17], and pyrolytic acid [18]. The utilization of pyrolysis liquid is already of interest in many areas such as biocide, insecticide, fungicide, plant growth stimulator, and source of valuable chemicals [13,18,19]. A high concentration of acetic acid and furfural make pyrolysis liquid a potent natural pesticide [12]. Due to its large composition of bioactive chemicals, such as organic acids, phenols, ketones, aldehydes, furans, and guaiacols, pyrolysis liquid was presumed to protect the wood against fungi and termite [15]. Barbero-López et al. [14] reported that pyrolysis distillates from spruce bark, birch bark, and fiber hemp at 275°C–350°C temperatures provide inhibition to the growth of wood-decaying fungi. Pyrolysis liquid made from rubberwood and bamboo at 400°C showed better anti-fungal activity towards white-rot, brown-rot, and sapstain fungi [20]. Oramahi et al. [17] also studied the pyrolysis liquid from oil palm trunk at a 350°C as an anti-termite.

Research on the pyrolysis of sugarcane bagasse has also been previously conducted, mainly on the pyrolysis parameters and products' characterization (biochar, bio-oil, and gas) [7,21–23]. Past studies reported that the optimum parameters to produce a high char yield were at 2°C/min–16°C/min of heating rate and at a temperature above 240°C, where significant differences of mass loss occurred at a temperature between 300°C and 500°C [7,22]. From those results, we could establish the slow pyrolysis experiments at a lower heating rate (1 °C/min and 10 °C/min) and intermediate temperature (400°C and 500°C), since the good quality of char could be obtained with higher temperatures [24].

Therefore, in this study, we aimed to evaluate the potential of sugarcane bagasse valorization via slow pyrolysis. The char resulting from different pyrolysis settings was assessed according to its yield and thermal properties, including the proximate, ultimate, and calorific analysis for evaluating its potential as an energy source. The chemical composition of pyrolysis liquids selected from the process, along with the test against the wood-decaying fungi and subterranean termite, were also investigated.

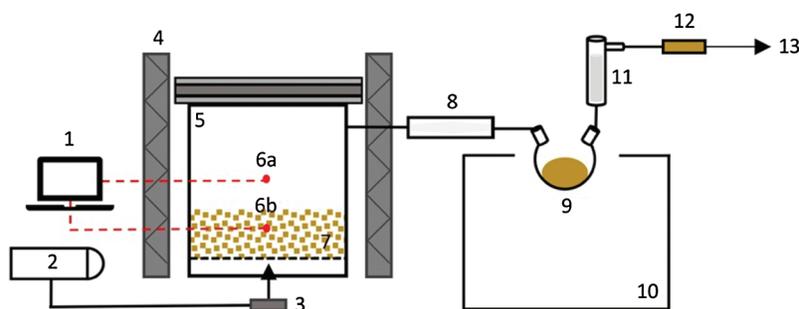
## 2 Material and Methods

### 2.1 Biomass Feedstock and Preparation

The sugarcane bagasse (*Saccharum* spp.) was obtained from three consecutive annual harvests from la Réunion Island (French overseas territory). The bagasse has a moisture content of 8.48% and was provided as a grounded form with a particle size of <2 mm. Before the pyrolysis process, the samples were oven-dried at a temperature of 105°C for 2 h to remove their moisture content (<1%). Samples of 100 g of bagasse were used for each pyrolysis experiment.

### 2.2 Slow Pyrolysis Set-up and Parameters

A bench-scale fixed bed reactor equipped with a set of condensation system was used for the slow pyrolysis experiments (Fig. 1). This device was placed in a muffle furnace oven, which generates heat to the reactor. The reactor vessel has a cylinder form with a 24 cm diameter and a height of 20 cm. The reactor cap has a 59.5 cm long connecting pipe, which allows smoke to exit the oven towards the condensation system. The condenser was circulated with the ice-cooled liquid, a mixture of water and ethylene glycol contained in the cooling bath. The condenser's temperature was held at the temperature of -15°C at the beginning of the pyrolysis experiments.



**Figure 1:** Experimental set-up of the pyrolysis reactor: (1) temperature control, (2) nitrogen gas supply, (3) flow control, (4) muffle furnace, (5) pyrolysis reactor, (6a and 6b) thermocouples, (7) bagasse/char, (8) condenser, (9) pyrolysis liquid, (10) cooling bath, (11) electrostatic precipitator, (12) silica gel, and (13) gas vent

The experiments were carried out at a temperature of 400°C and 500°C, a heating rate of 1°C/min and 10°C/min, and holding time of 30 min and 60 min, resulting in eight pyrolysis treatments in total. Two thermocouples were set up inside the reactor to monitor the evolution of temperature as a function of time to ensure the expected pyrolysis parameters. The constant flow of inert gas (N<sub>2</sub>) was applied at a 1 L/min rate, which removed the organic vapors and gas products from the reactor. The pyrolysis vapors were cooled down in a single-stage consisting of a double tube condenser maintained between 0°C and 5°C and an electrostatic precipitator. A bottle filled with silica gel was set up at the end of the condenser system, and the mass of the light-end compounds trapped here was used for mass balance purposes only. The yield of char and pyrolysis liquid was calculated based on the ratio between the final products' weight to biomass's initial weight. The pyrolysis liquid was then collected from the round flask, as

illustrated in Fig. 1. The gaseous product was not recovered in this study. Each pyrolysis experiment was repeated three times.

### 2.3 Bagasse and Char Analysis

The moisture content of bagasse [25] and char [26] were determined by heating the samples at 105°C until reaching a constant weight. For the ash content, the samples of bagasse [27] and char [26] were heated for 2 h (or until reaching stable weight) to 550°C and 710°C ± 10°C, respectively. The volatile matter of bagasse [28] and char [26] were measured by heating the samples to 900°C for 7 min. The fixed carbon was determined by difference. For the higher heating value (HHV) of biomass and char, the sample was experimentally measured using a bomb calorimeter PARR 6200 [29] and determined on its anhydrous basis (Eq. (1)).

$$HHV = 4.1855 \times \frac{E_{cal} \times \theta - Q_{fuse} - Q_N}{m} \times \frac{100}{100 - MC} \quad (1)$$

where 4.1855 is the conversion factor (from Cal to J),  $m$  is the mass of the sample (g),  $E_{cal}$  is the average value of the effective heat capacity of the calorimeter as determined during calibration tests (Cal/°C),  $\theta$  is the corrected temperature increase (°C),  $Q_{fuse}$  is the energy input from the combustion of the firing wire (Cal),  $Q_N$  is the energy input from the formation of nitric acid (from water in the liquid state, nitrogen and oxygen in the gaseous state) (Cal), and  $MC$  is the moisture content of the sample (%).

Lastly, the elemental analysis of C, H, and N content of biomass and char was performed using an elemental analyzer VariMACROcube referred to ASTM D5373 [30] and ISO 16948 [31]. The O/C and H/C atomic ratio were then determined from the obtained weight percent and atomic weights of the corresponding elements. The ash recovery (AR) was calculated following the study conducted by Ghysels et al. [10] and determined from the ash content of bagasse ( $Ash_{bag}$ ), the ash content of char ( $Ash_{char}$ ), and the char yield (CY) (Eq. (2)).

$$AR = CY \times \frac{Ash_{char}}{Ash_{bag}} \quad (2)$$

The energy yield (EY) and energy density (ED) were determined according to Yang et al. [32]. EY was expressed as the ratio of the heating value of char ( $HHV_{char}$ ) and bagasse ( $HHV_{bag}$ ) multiplied by the char yield (CY) (Eq. (3)). Meanwhile, ED was determined as the ratio of energy yield and char yield (Eq. (4)).

$$EY = CY \times \frac{HHV_{char}}{HHV_{bag}} \quad (3)$$

$$ED = \frac{EY}{CY} \quad (4)$$

### 2.4 Chemical Characterization and Water Content of Pyrolysis Liquid

Fourier transform infrared spectroscopy (FTIR) spectra of the pyrolysis liquid were recorded using a Perkin Elmer Frontier spectrometer equipped with an Attenuated Total Reflection (ATR). FTIR spectra were obtained at a nominal resolution of 4 cm<sup>-1</sup> for 4 scans in the range of 4000–650 cm<sup>-1</sup>. The spectra were treated using RStudio using the package of MALDIquantForeign [33,34].

Gas chromatography-mass spectrometry (GC-MS) analysis of pyrolysis liquid was done using an Agilent 6890N gas chromatograph and Agilent 5975 mass spectrometer, all materials and chemicals being provided by Agilent. A 1 mL of pyrolysis liquid sample was dissolved in 10 mL of analytical grade acetone and filtered with a nylon microfilter of 0.45 mm. Afterward, 1 mL of sample volume was

transferred into a glass vial followed by the introduction of internal standards (acetic-acid-d4, toluene-d8, phenol-d6, and phenanthrene-d10). The gas chromatograph was equipped with an electronically controlled split/splitless injection port and a capillary column DB-1701. The injection (1  $\mu$ L) was performed at 250°C in the split mode (split = 1/10). Helium was used as carrier gas at constant flow (1.9 mL/min). GC oven temperature was programmed to hold at 45°C for 4 min, then heated to 120°C at a rate of 3 °C/min, and finally increased to 270°C at a rate of 20 °C/min. The internal standards used for the quantification were acetic-acid-d4 and toluene-d8. For the splitless mode, the injection was performed at 250°C. The oven temperature program was as follows: the initial temperature was set at 45°C for 4 min, then increased to 250°C at a rate of 3 °C/min, increased again to 270°C at a rate of 20 °C/min and held there for 60 min. Phenol-d6 and phenanthrene-d10 internal standards were used for the quantification of the compounds. The mass spectrometer was fixed at 70 eV ionization energy (ion source temperature = 230°C). Detail apparatus specification and description of the GC–MS method have been previously described by Lê Thành et al. [35]. In addition, a dilution to 1/10 (v/v) of a mixture sample in acetone was also performed to obtain the actual concentration of some calibrated compounds that were higher than the upper limit of the calibration range.

The water content of pyrolysis liquid was determined using a Karl Fisher volumetric titration (Mettler Toledo, Karl Fisher V20) following the standard test method ASTM E 203-08 [36]. The pH of pyrolysis liquid was also measured using a pH meter (Portames<sup>®</sup> 911pH). Only pyrolysis liquid produced at a temperature of 400°C and 500°C, a heating rate of 10 °C/min, and a holding time of 60 min were analyzed.

### 2.5 Fungal Growth Inhibition Test

Pyrolysis liquid produced at a temperature of 400°C and 500°C, a heating rate of 10°C/min, and a holding time of 60 min were tested against two types of wood-decaying fungi: *Trametes versicolor* (white-rot, strain CTB 863 A) and *Coniophora puteana* (brown-rot, strain BAM Ebw. 15). The test was carried using a malt-agar medium in a 9-cm diameter Petri-dish. The malt-agar medium was prepared by mixing 40 g of malt (Quaron) and 20 g of agar (BioMérieux) in 1 L of deionized water and sterilized using an autoclave for 20 min at 121°C. To determine the minimum inhibitory concentration (MIC), various concentrations of pyrolysis liquid (v/v) were tested (0.05%, 0.10%, 0.15%, 0.20%, and 0.25%) by diluting them directly in the malt agar solution, right after its distribution (40 mL) into a Petri-dish. The mixture was then hand-stirred immediately until homogenized and let stay until it solidified. The two types of fungi, in the form of a 0.5 cm diameter mycelium disk, were then placed centrally on the Petri-dishes, with four dishes for each type of fungi and each concentration. Afterward, the Petri-dishes were incubated at 22°C and 70% Relative Humidity (RH). The inhibition rate ( $I_r$  in %) was determined by comparing the diameter of fungal growth of the treated medium with the control (Petri-dishes contained only malt-agar medium) following the method described by Kartal et al. [16] (Eq. (5)).

$$I_r = 1 - \frac{D_t}{D_c} \quad (5)$$

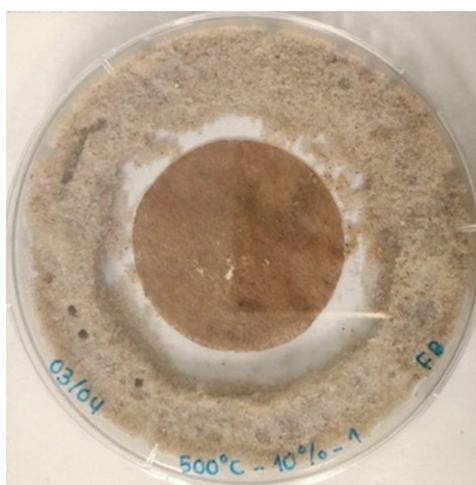
where  $D_c$  is the colony diameter of mycelium from the control Petri-dishes (mm), and  $D_t$  is the colony diameter (mm) of mycelium from tested the Petri-dishes containing pyrolysis liquid.

### 2.6 Termites Non-Choice Test

Termites test was evaluated using the pyrolysis liquid produced at the temperature of 400°C and 500°C, a heating rate of 10 °C/min, and a holding time of 60 min. A non-choice test was conducted against *Reticulitermes flavipes* (ex. *santonensis*) termites using a Joseph filter paper (grammage 25 g/m<sup>2</sup>, Filtratech) made of pure cellulose (diameter of 4.25 cm). Pyrolysis liquid was diluted in ethanol (v/v) (Honeywell) at a 5% and 10% concentration and treated to the filter paper by immersing them for about

60 s. Two controls were also set up by soaking the filter paper in deionized water and ethanol, separately. Six replicates were tested for each treatment. All of the treated filter papers were air-dried before the test for at least 10 min or 20 min for paper treated with water.

A 9-cm diameter of Petri-dish was filled with 30 g of wet Fontainebleau sand (4 vol of sand/1 vol of deionized water). The Joseph papers were placed on the plastic mesh to avoid water saturation. The test device is illustrated in Fig. 2. Then, a total of 20 termite workers were introduced into each petri dish and put in a dark climatic chamber at 27°C and RH > 75% for 4 weeks, watered and observed regularly. The termite mortality and the loss of paper area were determined at the end of the test. A simple tool was developed to calculate paper area loss using ImageJ software and Trainable Weka Segmentation (TWS), a machine learning tool for rapidly measuring images [37].



**Figure 2:** The test device for termite's evaluation using a filter paper

### 2.7 Data Analysis

Multivariate regression analysis was conducted to determine each parameter's correlation and effects on the yield values and char properties. Temperature, heating rate, and holding time was the independent parameters. We then determined the significance of each of those parameters individually using  $p$ -value and globally by seeing their  $R^2$  value.

For the anti-fungal and anti-termites test, a  $t$ -test analysis was used to compare the data from each specimen and treatment. The data were analyzed in order to see their groupings at a significance level of  $p < 0.05$ . The analysis was conducted using Microsoft Excel, R [38], and RStudio [34].

## 3 Results and Discussion

### 3.1 Biomass Characterization

Tab. 1 summarizes the proximate, ultimate, and calorific analysis of sugarcane bagasse. The proximate analysis provides information related to the burning quality of biomass [39]. Sugarcane bagasse contains 8.5% of moisture content, 82.4% of volatile matter, 13.4% of fixed carbon, and relatively higher ash (4.2%) compared to other studies (see Tab. 1). The ash content of biomass tends to be higher when smaller particles were used [7,21]. Ash content has a detrimental effect as it is an inorganic, incombustible material, and thus can lower the burning rate. The amount of ash content depends on the nature of biomass. As an illustration, rice straw has 14% of ash content, while wood biomass such as birch wood has the lowest ash content, which is approximately 0.3% [12,32]. Sugarcane bagasse also

contains a very low fuel ratio (fixed carbon/volatile matter) due to the high content of volatile matter. A fuel ratio of 1 or less is considered very low, while, as a comparison, the anthracite has a fuel ratio of 12 [40].

**Table 1:** Proximate and ultimate analysis of sugarcane bagasse

Characteristics	Sugarcane bagasse	Carrier et al. [7] <sup>a</sup>	Garcia-Perez et al. [21] <sup>b</sup>	Varma and Mondal [39] <sup>c</sup>
<i>Proximate analysis (dry wt.%)</i>				
-Moisture	8.5	15.4	6.0	5.4
-Volatile matter (VM)	82.4	–	82.1	80.2
-Ash content	4.2	3.1	1.6	3.1
-Fixed carbon (FC)	13.4	–	16.3	11.3
Fuel ratio (FC/VM)	0.16	–	0.20	0.14
<i>Ultimate analysis (dry wt.%)</i>				
-Carbon (C)	47.3	50.2	49.6	44.8
-Hydrogen (H)	5.7	5.6	6.0	5.8
-Nitrogen (N)	0.4	1.1	0.5	0.2
-Oxygen + Sulphur (O + S)	46.7	40.0	43.8	49.0
HHV (MJ/kg)	18.1	18.5	18.5	18.0

<sup>a</sup> Sugarcane bagasse obtained from South Africa.

<sup>b</sup> Sugarcane bagasse obtained from Florida, United States.

<sup>c</sup> Sugarcane bagasse obtained from India.

The ultimate analysis demonstrates the elemental composition of biomass. Sugarcane bagasse has 47.3% of carbon, 5.7% of hydrogen, 0.4% of nitrogen, and 46.7% of total oxygen and sulfur. The carbon and hydrogen contents are good criterion for power generation and contributes most to its calorific value [41]. The HHV of sugarcane bagasse was 18.1 MJ/kg. It is one of the most important energy analysis parameters, defined as the released energy per unit mass or volume from the material after the complete combustion. In contrast, the presence of oxygen decreases the fuel's calorific value, while the low content of nitrogen and sulfur is imperative in terms of environmental aspects [42]. The analysis demonstrates that the present biomass is comparable to other sugarcane bagasse reported in different studies (Tab. 1) and possesses the potential as pyrolysis feedstock.

### 3.2 Effect of Pyrolysis Parameters to the Product Yields

The evolution of temperature as a function of time was monitored and corresponded well with the expected parameters. The yield of char and pyrolysis liquid is presented in Tab. 2, followed by their coefficient of regression given in Tab. 3. The temperature was observed as the most important parameter, influencing the char yield ( $p$ -value =  $5.37 \times 10^{-12}$ ) and liquid yield ( $p$ -value =  $2.41 \times 10^{-5}$ ). The two heating rates (1 °C/min and 10 °C/min) also showed a significant effect on char yield ( $p$ -value =  $5.33 \times 10^{-6}$ ) and liquid yield ( $p$ -value =  $1.18 \times 10^{-5}$ ). Meanwhile, there is no significant difference in the product yields for the holding time of 30 min and 60 min.

**Table 2:** Energy analysis of bagasse char and the yield of bagasse char and pyrolysis liquid

Process condition	Char yield (%)	HHV (MJ/kg)	EY (%)	ED	Liquid yield (%)
400°C; 1 °C/min; 30 min	35.12 ± 0.96	27.44 ± 0.03	53.12	1.51	49.67 ± 0.28
400°C; 1 °C/min; 60 min	34.63 ± 0.20	27.41 ± 0.14	52.33	1.51	49.18 ± 1.33
400°C; 10 °C/min; 30 min	33.26 ± 0.60	27.20 ± 0.05	49.87	1.50	52.91 ± 0.60
400°C; 10 °C/min; 60 min	33.78 ± 0.60	26.86 ± 0.14	50.01	1.48	51.40 ± 1.12
500°C; 1 °C/min; 30 min	31.32 ± 0.25	28.52 ± 0.04	49.25	1.57	51.24 ± 0.71
500°C; 1 °C/min; 60 min	30.42 ± 0.31	28.67 ± 0.09	48.08	1.58	52.61 ± 0.95
500°C; 10 °C/min; 30 min	28.81 ± 0.19	27.81 ± 0.06	44.17	1.53	54.08 ± 0.50
500°C; 10 °C/min; 60 min	28.97 ± 0.36	27.85 ± 0.09	44.47	1.54	55.46 ± 0.22

**Table 3:** Coefficient of regression between pyrolysis parameters and the yield of bagasse char and pyrolysis liquid based on multivariate analysis

Dependent variable	Independent variable	Estimate	Std. error	t-value	Pr (> t )	R <sup>2</sup>
Char Yield	Intercept	52.883	1.182	44.726	3.10 × 10 <sup>-18</sup>	0.958
	Temperature	-0.044	0.002	-17.874	5.37 × 10 <sup>-12*</sup>	
	Heating rate	-0.186	0.028	-6.676	5.33 × 10 <sup>-6*</sup>	
	Holding time	-0.003	0.008	-0.306	7.63 × 10 <sup>-1</sup>	
Liquid Yield	Intercept	38.611	2.112	18.280	3.81 × 10 <sup>-12</sup>	0.821
	Temperature	0.026	0.004	5.861	2.41 × 10 <sup>-5*</sup>	
	Heating rate	0.309	0.050	6.242	1.18 × 10 <sup>-5*</sup>	
	Holding time	0.005	0.015	0.314	7.58 × 10 <sup>-1</sup>	

\*Correlation is significant at confidence level 0.05.

Lower temperature promotes a higher char yield; however, it reduces the liquid yield and vice versa. The char yield was decreased from the temperature of 400°C–500°C as much as 12.65% on average. Katyal et al. [22] stated that the significant conversion of bagasse in the slow pyrolysis process occurs between 300°C and 500°C, while at the higher temperature (500°C–700°C), the decrease of char yield was considered insignificant (3%–4%). In Tab. 2, the ED of char was found higher (1.53–1.58) at higher pyrolysis temperatures, as the HHV also increases as the temperature increase. From all the conditions tested, the average of char and liquid yield varied between 28.81%–35.12%, and 49.67%–55.46%, separately. Aboulkas et al. [43] reported that pyrolysis liquid made from algal waste reaches a maximum yield at the temperature range of 500°C–550°C, while it decreases at the temperature of 600°C. At higher temperatures, the secondary reaction is likely to occur and produce a significant amount of gas.

Results showed that the heating rate of 10°C/min increased the liquid yield, but decreased the char yield. Further, Katyal et al. [22] reported that the effect of heating rate on the char yield is more noticeable at temperatures lower than 400°C. It conveys a negligible effect of the heating rate at a higher temperature when char production is the main target. The reduction of char yield at a higher heating rate has been explained as the result of the secondary cracking of the pyrolysis vapors and secondary decomposition of the char [7]. While preferring the high liquid yield, the higher heating rate should be chosen. However, at pyrolysis of 500°C, Aboulkas et al. [43] reported no significant impact on the char or liquid yield using a

higher heating rate (10°C/min–50°C/min). In our study, the highest liquid yield was observed at a temperature of 500°C and a heating rate of 10°C/min (55.46%).

The decomposition of biomass during pyrolysis is associated with cellulose, hemicellulose, and lignin content. Varma and Mondal [39] reported that sugarcane bagasse comprises 47.6% of cellulose, 39% of hemicellulose, and a considerable low content of lignin (11.2%). According to Morais et al. [44], who investigated the thermal behavior of sugarcane bagasse, there are at least three critical stages in this process: (1) Dehydration, (2) Degradation of hemicellulose and cellulose, and (3) Lignin decomposition. They investigated the maximum loss rate of hemicellulose was observed at 250°C, cellulose at 330°C, and lignin between 190°C and 500°C, with the maximum mass loss rate occurs at 430°C. Hemicellulose and amorphous part of cellulose principally yield the condensable volatiles, whereas cellulose and lignin's degradation produces gases, tar, and char [22].

### 3.3 Effect of Pyrolysis Parameters to the Bagasse Char Properties

The proximate and ultimate analysis of bagasse char from different pyrolysis conditions is presented in Tab. 4, followed by their coefficient of regression illustrated in Tab. 5. Temperature and heating rate were found to significantly influence the content of volatile, ash, fixed carbon, elemental carbon, and HHV, while holding time contributes to the volatile matter content only. After being pyrolyzed, bagasse char has a volatile matter of 24.31%–25.80% at 400°C, and 13.01%–16.43% at 500°C. Previous studies confirmed that the volatile matter is affected by the pyrolysis temperature [22,45], as the decrease of volatile matter follows the increase of temperature. The reduction of volatile content between the two temperatures occurs due to the cellulose and lignin decomposition. As the volatile released, fixed carbon and ash content percentage increased. The highest fixed carbon was observed at 68.50%–73.01%, while the ash content was found to be higher than 10%. The ash content in the present study varied from 10.79%–15.50%; however, it is still lower than the study conducted by Varma and Mondal [39], who found it to be 16.25%. High ash recoveries presented in Tab. 4 also showed that ash is retained within the char and attractive for its application as soil amendments [10]. Higher temperature and lower heating rate also favor the higher carbon content and higher HHV. At 500°C, lower heating resulted in higher carbon content (73.53%–75.70%) than those obtained at the higher heating rate (71.95%–72.90%). HHV produced from the sugarcane bagasse pyrolysis also increases from 18.14 MJ/kg to 28.67 MJ/kg at the higher temperature and lower heating rate. The char has a higher carbon content and a lower oxygen content compared to its raw material. The O/C and H/C molar ratio of char were lower than sugarcane bagasse due to the loss of H and O during pyrolysis. The O/C and H/C molar ratio of sugarcane bagasse was found as 0.74 and 1.43, respectively.

**Table 4:** Proximate and ultimate analysis of bagasse char at different pyrolysis conditions

Process conditions	Proximate analysis (wt.%)			Ash recovery (wt.%)	Ultimate analysis (wt.%)				O/C ratio	H/C ratio
	Volatile matter	Ash content	Fixed carbon		C	H	N	O		
400°C; 1 °C/ min; 30 min	25.67 ± 1.15	10.79 ± 0.12	63.54	86.63	69.50 ± 0.14	3.81 ± 0.03	0.55 ± 0.03	26.14	0.28	0.66
400°C; 1 °C/ min; 60 min	24.31 ± 0.58	11.72 ± 0.78	63.97	88.81	70.23 ± 0.90	3.66 ± 0.04	0.58 ± 0.05	25.53	0.27	0.63
400°C; 10 °C/ min; 30 min	25.80 ± 0.06	12.81 ± 0.26	61.39	95.86	70.35 ± 0.07	3.32 ± 0.00	0.65 ± 0.02	25.68	0.27	0.57
400°C; 10 °C/ min; 60 min	24.56 ± 0.54	13.15 ± 0.51	62.29	97.91	69.40 ± 0.42	3.46 ± 0.07	0.62 ± 0.01	26.52	0.29	0.60

(Continued)

**Table 4 (continued).**

Process conditions	Proximate analysis (wt.%)			Ash recovery (wt.%)	Ultimate analysis (wt.%)				O/C ratio	H/C ratio
	Volatile matter	Ash content	Fixed carbon		C	H	N	O		
500°C; 1 °C/min; 30 min	14.33 ± 0.21	13.00 ± 0.85	72.67	87.80	73.53 ± 0.40	3.07 ± 0.00	0.53 ± 0.04	22.87	0.23	0.50
500°C; 1 °C/min; 60 min	13.01 ± 0.31	13.99 ± 0.49	73.01	94.99	75.70 ± 0.00	3.08 ± 0.04	0.61 ± 0.00	20.61	0.20	0.49
500°C; 10 °C/min; 30 min	16.01 ± 0.36	15.50 ± 0.79	68.50	98.58	72.90 ± 0.57	2.74 ± 0.03	0.62 ± 0.03	23.74	0.24	0.45
500°C; 10 °C/min; 60 min	16.43 ± 0.06	13.07 ± 0.57	70.50	84.52	71.95 ± 0.78	2.85 ± 0.01	0.57 ± 0.06	24.63	0.26	0.48

**Table 5:** Coefficient of regression between pyrolysis parameters and the quality of char based on multivariate analysis

Dependent variable	Independent variable	Estimate	std. error	t-value	Pr (> t )	R <sup>2</sup>
Volatile matter content	Intercept	67.055	1.918	34.953	2.83 × 10 <sup>-17</sup>	0.976
	Temperature	-0.103	0.004	-25.892	4.24 × 10 <sup>-15*</sup>	
	Heating rate	0.139	0.044	3.143	5.93 × 10 <sup>-3*</sup>	
	Holding time	-0.036	0.013	-2.711	1.48 × 10 <sup>-2*</sup>	
Ash content	Intercept	3.464	1.876	1.847	8.04 × 10 <sup>-2</sup>	0.657
	Temperature	0.019	0.004	4.973	8.44 × 10 <sup>-5*</sup>	
	Heating rate	0.156	0.043	3.6428	1.73 × 10 <sup>-3*</sup>	
	Holding time	0.003	0.013	0.260	7.98 × 10 <sup>-1</sup>	
Carbon content	Intercept	55.417	2.379	23.296	3.42 × 10 <sup>-13</sup>	0.806
	Temperature	0.037	0.005	7.677	1.43 × 10 <sup>-6*</sup>	
	Heating rate	-0.111	0.053	-2.089	5.42 × 10 <sup>-2*</sup>	
	Holding time	0.008	0.016	0.496	6.27 × 10 <sup>-1</sup>	
HHV	Intercept	23.883	0.350	68.144	3.96 × 10 <sup>-26</sup>	0.935
	Temperature	0.010	0.001	13.870	4.83 × 10 <sup>-12*</sup>	
	Heating rate	-0.065	0.008	-8.463	3.30 × 10 <sup>-8*</sup>	
	Holding time	-0.002	0.002	-0.967	3.44 × 10 <sup>-1</sup>	

\*Correlation is significant at confidence level 0.05.

For energy application, low volatile content, low ash content, high fixed carbon, high carbon content, high H/C, low O/C, and high HHV should be preferred. Meanwhile, high ash recovery and high stable carbon contents (from fixed carbon) can be a good indicator of soil amendments. Generally speaking, bagasse char obtained at 500°C, and a heating rate of 1 °C/min or 10 °C/min has met all the criteria. Indeed, a lower heating rate promotes higher char yield and higher HHV. However, in the present study, the HHV of char from the pyrolysis of 500°C and the heating rate of 10 °C/min (27.85 MJ/kg) is already

higher or comparable to the HHV of char produced on the same temperature and heating rate  $<10$  °C/min by other studies [7,46]. Carrier et al. [7] found that the optimum HHV of bagasse slow pyrolysis (27.67 MJ/kg) was obtained at  $T > 450$ °C and a heating rate lower than 8 °C/min. As a heating rate of 10 °C/min also promotes higher pyrolysis liquid yield, it is reasonable to choose a higher temperature and higher heating rate in this study to co-produce the bagasse char and bagasse pyrolysis liquid. Since the temperature was found as a dominant factor influencing the char properties, char yield, and liquid yield, we evaluated the two types of pyrolysis liquid produced as a candidate for anti-fungal and anti-termite agents.

### 3.4 Characterization of Pyrolysis Liquid

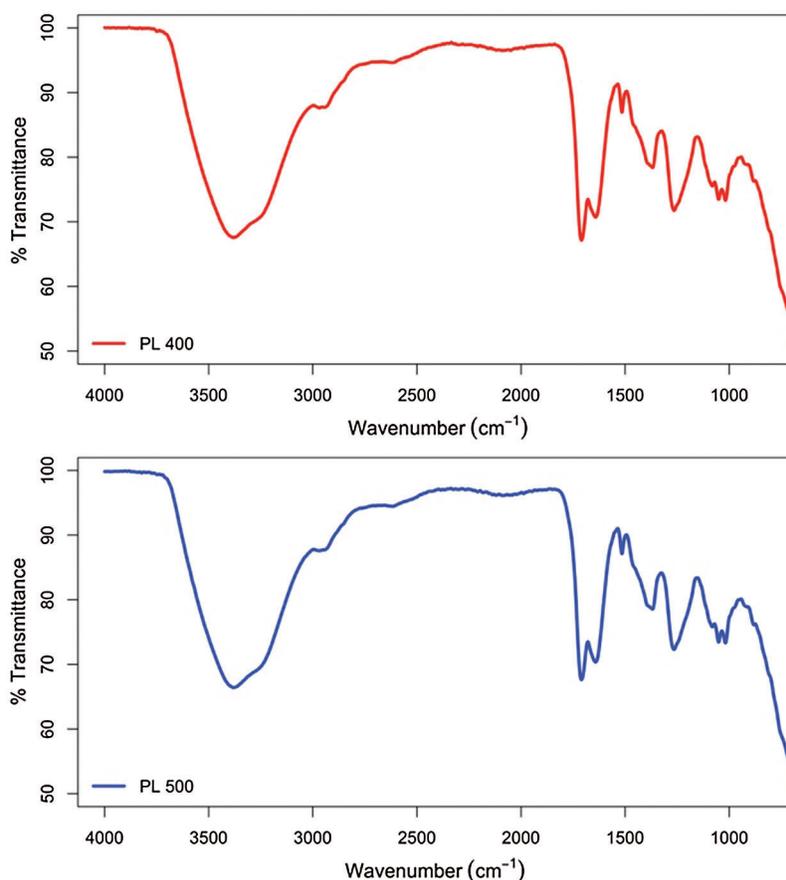
Pyrolysis liquid was obtained in a dark brown color with a pungent, smoky odor. Previously, Demirbas [47] has stated that the pyrolysis liquid consisted of two distinct parts: an aqueous phase consisting of a mixture of light oxygenates; and tar, a non-aqueous phase containing insoluble organics of high molecular weight. At the time of experiment, the liquid obtained was not entirely in a separated phase (the soluble tar was presented and mixed with the aqueous phase). The heavy tar was presented, and it can be separated from the aqueous phase by the decantation [12]. However, not all the heavy tar was successfully recovered at the end of pyrolysis experiments. Some of them were trapped in the condensation system; however, this amount is not significant (~1%). To avoid the ambiguity, the term of pyrolysis liquid used in this study represents the total condensed liquid obtained from the pyrolysis experiments (crude pyrolysis liquid).

Fig. 3 demonstrated the FTIR spectra, which represents the functional group of the pyrolysis liquid. According to the spectra obtained, pyrolysis liquid was confirmed to have various chemical components, including oxygenated compounds, such as carboxylic acids, alcohols, phenols, aldehydes and ketones, and other chemical components such as alkanes, alkenes, and aromatics. The spectra of pyrolysis liquid obtained from the slow pyrolysis at the temperature of 400°C and 500°C were similar. The peak between  $3400\text{ cm}^{-1}$  and  $3200\text{ cm}^{-1}$  presents the O–H stretching vibration corresponding to alcohols and phenols' content. The weak peaks observed between  $3000\text{ cm}^{-1}$  and  $2800\text{ cm}^{-1}$  (C–H stretching) imply the presence of alkanes. The C=O stretching vibrations between  $1780\text{ cm}^{-1}$  and  $1650\text{ cm}^{-1}$  indicate the presence of ketones, aldehydes, and carboxylic acids [48,49]. Peaks at  $1513\text{ cm}^{-1}$  (C=C stretching vibration) correspond to the presence of alkenes and aromatics. Peaks at around  $1365\text{ cm}^{-1}$  signify the presence of alkane due to the C–H vibrations [50]. The region between  $1300\text{ cm}^{-1}$  and  $900\text{ cm}^{-1}$  (C–O stretching and O–H deformation vibration) was reported to correspond to the presence of alcohols, carboxylic acids, ethers, esters, and phenols, while for the region between  $900\text{ cm}^{-1}$  and  $700\text{ cm}^{-1}$  (C–H bending) corresponds to the existence of aromatic compounds [39]. The spectra obtained were in accordance with the study carried out by Varma and Mondal [39].

The main organic compounds present in the organic fraction of pyrolysis liquid are given in Tab. 6. Chemical analysis using a GC–MS allowed the quantification of 56 chemical compounds, where the total organic contents quantified were 29.63% and 25.63% for pyrolysis liquid produced at 400°C and 500°C, respectively. Indeed, the pyrolysis liquid contained some non-volatile compounds that are not GC-eluted, in which the non-quantified fraction represented 28%–32% of the liquid products. The said compounds are not determined in this study.

The chemical composition of the pyrolysis liquid is influenced by the content of cellulose, hemicellulose, and lignin in biomass. The holocellulose generated furans and carbohydrates, whereas the lignin led to the formation of phenolic compounds [47]. According to the analysis, pyrolysis liquid consisted of carboxylic acids, alcohols, aldehydes, ketones, furans, and anhydrosugars, whereas the acetic acid, glycolaldehyde, 1-hydroxy-2-propanone, methanol, formic acid, and levoglucosan were the principal compounds. Excluding the water, acetic acid was the significant compound composing the pyrolysis liquid, presenting 7.22%–8.48% of the liquid. Sugarcane bagasse is reported to be high in the content of

xylose (15.5%–28.9%) [51], which contribute to the formation of acetic acid by the elimination of acetyl groups from the xylose unit [15,47]. Demirbas et al. [47] studied the mechanism of the pyrolysis reaction of biomass. For example, methanol resulted from the methoxyl groups of uronic acid, whereas furfural came from the xylose dehydration. Additionally, the degradation of xylan produced water, methanol, 1-hydroxy-2-propanone, 1-hydroxy-2-butanone, 2-furfuraldehyde, acetic, propionic, and formic acid.



**Figure 3:** FTIR spectra of pyrolysis liquid (PL) at the temperature of 400°C and 500°C

**Table 6:** Main organic compound (%) presented in the organic fraction of pyrolysis liquid

Compound name	Pyrolysis liquid (%)		Classification
	400°C	500°C	
Acetic acid	8.48	7.22	Acid
Glycolaldehyde	3.35	3.18	Aldehyde
1-hydroxy-2-propanone	3.32	2.91	Ketone
Methanol	2.28	1.90	Alcohol
Levogluconan	1.79	1.75	Sugar
Formic acid	1.70	1.60	Acid

**Table 6 (continued).**

Compound name	Pyrolysis liquid (%)		Classification
	400°C	500°C	
Furfural	1.19	0.45	Furan
1-acetyloxy-2-propanone	0.87	0.78	Ketone
2-butanone,1-hydroxy-	0.82	0.73	Ketone
2-hydroxy-2-cyclopenten-1-one	0.66	0.61	Ketone
DGP*	0.63	0.52	Sugar
Acetaldehyde	0.48	0.39	Aldehyde
3-methyl-1,2-cyclopentanedione	0.41	0.38	Ketone
Propionic acid	0.37	0.31	Acid
Formaldehyde	0.33	0.26	Aldehyde
2,6-dimethoxyphenol	0.30	0.27	Guaiacol
2-methoxy-4-vinylphenol	0.24	0.18	Guaiacol
Phenol-2-methoxy	0.22	0.20	Guaiacol
Phenol	0.21	0.22	Phenol
Phenol-2-methoxy-4-methyl	0.16	0.14	Guaiacol
Phenol-4-ethyl-2-methoxy	0.15	0.12	Guaiacol
Phenol-3-methyl+Phenol-4-methyl	0.11	0.11	Phenol

\*DGP = 1,4:3,6-dianhydro- $\alpha$ -D-glucopyranose.

There is only a slight variation of the chemical compounds of pyrolysis liquid produced at the temperature 400°C and 500°C. However, higher organic content was found for the pyrolysis liquid produced at 400°C, which also possessed a higher content of furfural (1.19%). An increase in pyrolysis temperature contributed to the increase of phenolic compounds and decrease to the organic acid content [52]. Further, no polycyclic aromatic hydrocarbon compounds (PAHs) were found, or they were insignificant (less than 1 ppm). The PAHs content must be taken carefully into account because of their toxicity towards the environment and human health [53]. In contrast to creosote, pyrolysis liquid presented in this study could be considered to be more environmentally advantageous. According to the European standard EN 13991 [54], the maximum content of benzo(a)pyrene in creosote class B is 50 mg/kg. Cordella et al. [55] stated that PAHs are formed with a higher proportion at a higher temperature (650°C, 100°C/min). Additionally, the formation of phenol increases as the continuous secondary degradation occurs with the rising of temperature [47].

Pyrolysis liquid has  $42.34 \pm 0.24\%$  and  $42.30 \pm 0.14\%$  of water content at 400°C and 500°C temperatures, respectively. Temiz et al. [15] reported that bio-oil made from giant cane has a water content of 34%. Variation of water content is affected by the moisture content of biomass and dehydration reaction during pyrolysis. In addition, the condenser temperature, which is maintained at 0°C during the pyrolysis process also likely to contribute to the high levels of water, acids, alcohol, and ketones [56]. The pH was recorded at 2, in which the acidic characteristic was due to the presence of the organic acid, aldehydes, and phenols [39].

### 3.5 Effect of Pyrolysis Liquid in Inhibiting Fungi

Two types of pyrolysis liquid were revealed their efficacy to inhibit the growth of *C. puteana* and *T. versicolor* (Tab. 7). Fig. 4 showed the comparison of the fungal growth between the treated and untreated samples. Pyrolysis liquid produced at a temperature of 400°C slightly tends to be more active, delaying the fungal growth at low concentration, likely because of the relatively higher organic materials content. According to the GC-MS analysis, pyrolysis liquid at 400°C contains higher organic acid and furfural than the pyrolysis liquid obtained at 500°C. At a 0.15% concentration, pyrolysis liquid at 400°C caused a 100% inhibition of *C. puteana*. Also, to inhibit the fungal growth completely, pyrolysis liquid at a concentration of 0.20% and 0.25% was required to suppress the growth of *C. puteana* and *T. versicolor*, respectively. The different effect of pyrolysis liquid towards the fungi was likely correlated to the different metabolic rates or the enzyme released by the fungi [14].

**Table 7:** The percentage of inhibition of the pyrolysis liquid to the growth of *Coniophora puteana* and *Trametes versicolor*

Pyrolysis liquid	Concentration (%)	Inhibition (%)*	
		<i>C. puteana</i>	<i>T. versicolor</i>
400°C	0.05	20.71 ± 5.70 (b)	20.55 ± 2.91 (a)
	0.10	47.34 ± 2.70 (cd)	44.06 ± 1.64 (c)
	0.15	<b>100 ± 0.00 (e)</b>	64.23 ± 12.14 (d)
	0.20	<b>100 ± 0.00 (e)</b>	93.30 ± 13.39 (e)
	0.25	<b>100 ± 0.00 (e)</b>	<b>100 ± 0.00 (f)</b>
500°C	0.05	10.94 ± 3.51 (a)	21.44 ± 4.39 (a)
	0.10	37.70 ± 9.69 (c)	36.79 ± 1.51 (b)
	0.15	58.09 ± 7.41 (d)	57.52 ± 7.79 (d)
	0.20	<b>100 ± 0.00 (e)</b>	84.87 ± 17.55 (e)
	0.25	<b>100 ± 0.00 (e)</b>	<b>100 ± 0.00 (f)</b>

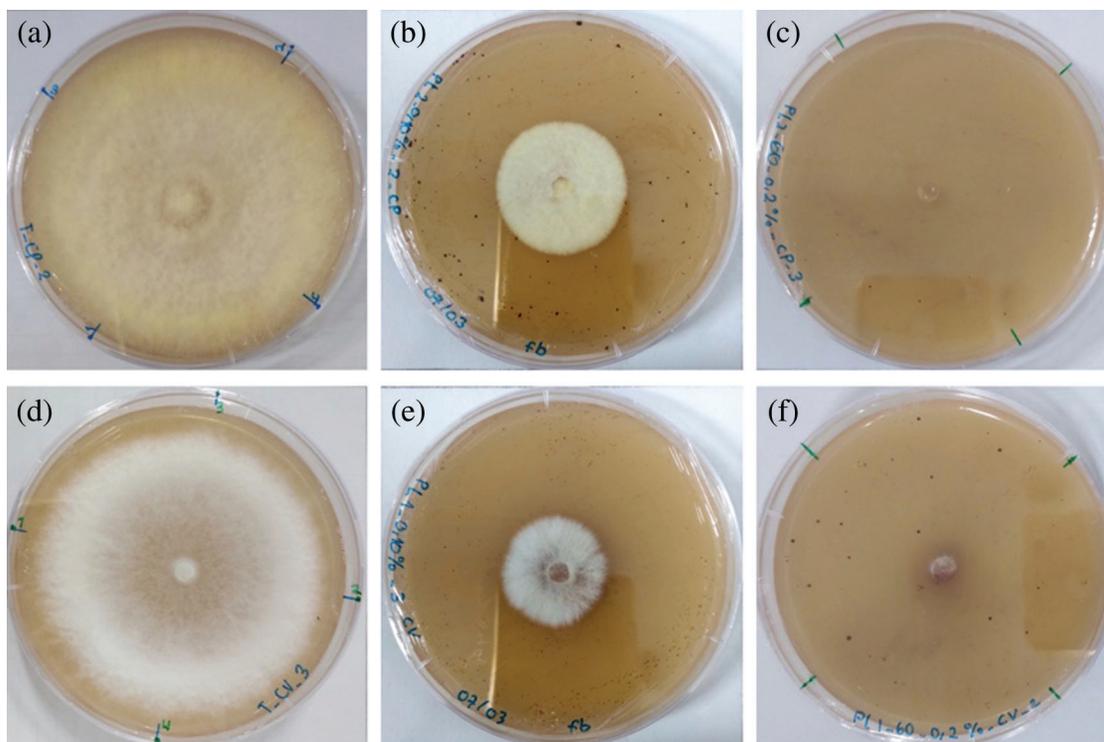
\*Means within each column and factor followed by the same letter are not significantly different.

\*Means (n = 4) ± SD.

It is difficult to compare the results with other studies since the high variability of pyrolysis liquid influenced by pyrolysis parameters and the nature of biomass used. However, mostly, at a concentration of less than 1%, pyrolysis liquid has an excellent inhibition activity against various types of wood-decaying fungi. Barbero-López et al. [14] found that the MIC values for the complete inhibition were between 0.5% and 1% by using the liquid product from the pyrolysis of tree bark and fiber hemp at the temperature of 275°C–350°C. Meanwhile, Kartal et al. [16] reported the tar oil made from the slow pyrolysis of macadamia nut shells effectively inhibited the brown-rot fungi, white-rot, and sap-staining fungi tested at the concentration of 0.20%.

The anti-fungal activity of pyrolysis liquid was likely the result of the synergetic effect from the various chemical compounds. The inhibition can also occur probably because of the compounds affect the enzymatic activity of the fungi. Several authors have reported that propionic acid [14], acetic acid [17], and furfural [12,18] exhibited anti-fungal activity. Barbero-López et al. [14] reported that at the concentration of 0.1% or below, propionic acid alone could completely inhibit fungal growth such as *C. puteana*. Guaiacols derivatives, which were also found in this study, such as phenol-2-methyl, phenol-3-methyl,

phenol-4-methyl, phenol-2-methoxy-4-methyl, and phenol-4-ethyl-2-methoxy were also responsible for the ability of pyrolysis liquid to inhibit fungi [57,58].



**Figure 4:** Growth of *C. puteana* (a = control; b = pyrolysis liquid at 0.10% concentration; c = pyrolysis liquid at 0.20% concentration) and *T. versicolor* (a = control; b = pyrolysis liquid at 0.10% concentration; c = pyrolysis liquid at 0.20% concentration)

### 3.6 Effect of Pyrolysis Liquid in Termite's Non-Choice Test

The effect of pyrolysis liquid on the paper area loss and mortality of *R. flavipes* are presented in Tab. 8. Control paper samples treated with deionized water and ethanol represent the equal portion of paper area loss (76.22% and 79.3%, respectively) with a low rate of mortality (8.33%). Meanwhile, paper treated with ethanol was proved not to affect the termites feeding activity. Fig. 5 showed the remaining area of paper consumed by termites. At a 5% concentration, termites consumed less portion of the paper (18.01%–28.15%), with a small number of termites survived during a 4-week test. Statistically, pyrolysis liquid obtained at 400°C is more effective when applied at a 5% concentration. However, using a 10% concentration, both pyrolysis liquids caused complete mortality, even if, in this non-choice test, termites still consumed a little portion of the paper (5.56%–7.03%). A higher rate of termite's mortality showed the toxic effect of pyrolysis liquid.

When termites were subjected to the treated paper on the first day of experimentation, termites seemed to stay away from the treated samples, probably because of the unpleasant odor from the pyrolysis liquid. It has been known that pyrolysis liquid has a smoky flavor characteristic due to the presence of ketones, phenol, guaiacol, and syringol [19,59]. The volatile compounds released during the experimentation was believed to repel the termites as reported by Oramahi et al. [60]. However, since the results showed that some termites have to consume the treated paper to die, the repellency is not enough, and the ingestion is necessary to make the termite die. Pyrolysis liquid produced from the biomass pyrolysis has been

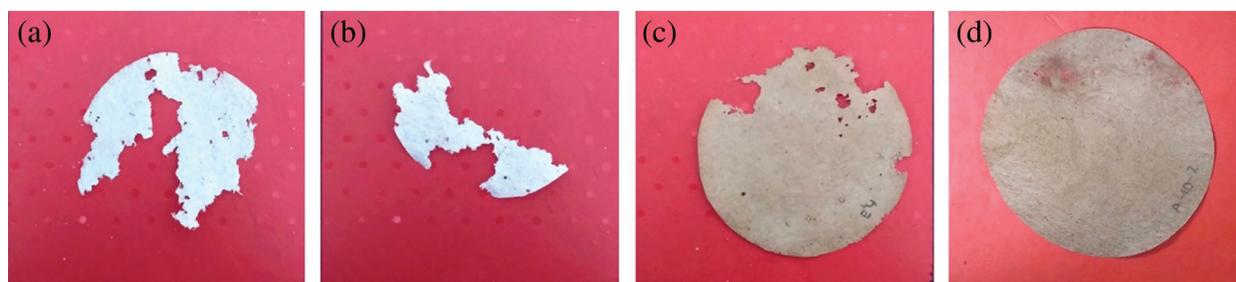
reported to have activity against *R. speratus*, *Coptotermes formosanus*, *C. curvignathus*, and *Odontotermes* sp [61]. The presence of formaldehyde in the pyrolysis liquid, akin to that of carbamate, can affect insects' nervous system [62]. Acetic acid and phenols also contribute to the termiticidal activity. Yatagai et al. [63] stated that formic acid showed strong termiticidal activities against *R. speratus*.

**Table 8:** Percentage of paper area loss and mortality of termites

Treatment	Paper area loss (%)*	Mortality (%)*
PL 400°C_5%	18.01 ± 3.11 (c)	93.33 ± 9.83 (c)
PL 500°C_5%	28.15 ± 6.01 (b)	89.17 ± 8.61 (b)
PL 400°C_10%	5.56 ± 2.75 (d)	<b>100.00 ± 0.00 (c)</b>
PL 500°C_10%	7.03 ± 3.41 (d)	<b>100.00 ± 0.00 (c)</b>
Control (ethanol)	79.30 ± 10.11 (a)	8.33 ± 5.16 (a)
Control (water)	76.22 ± 5.00 (a)	8.33 ± 6.06 (a)

\*Means within each column and factor followed by the same letter are not significantly different.

\*Means (n = 6) ± SD.



**Figure 5:** Paper consumed by termites (a = control (water); b = control (ethanol); c = pyrolysis liquid at 5% of concentration and d = pyrolysis liquid at 10% of concentration)

### 3.7 The Potential of Pyrolysis Liquid for Wood Protection

Pyrolysis liquid obtained from the slow pyrolysis of sugarcane bagasse was proven to be effective against fungi and termites. The effectiveness of pyrolysis liquid was believed to be the result of the synergetic action of its various chemicals, notably organic acid and phenolic compounds. The crude pyrolysis liquid obtained from this study has already provided good performances against fungi and termites, suggesting that the fractionation of the specific compound might be unnecessary, in accordance with the study by Mohan et al. [57]. Its potential utilization for wood protection has also already attracted some researchers [15,16] to conduct various studies. For instance, Lourençon et al. [64] reported that bio-oil from fast pyrolysis of residual eucalypt wood fines was an efficient anti-fungal and hydrophobic agent for wood. Barbero-López et al. [14] also reported, typically, the liquid products from slow pyrolysis show higher fungal activity than torrefaction-produced liquid. This is understandable because torrefaction, as a low-temperature pyrolysis process (between 225°C and 300°C), usually produces a liquid with lower organic compounds, mainly organic acid and water.

Indeed, current issues related to the environmental concerns and governmental regulation in banning toxic compounds of the wood preservatives led to the various efforts to search for alternatives. Different types of material such as essential oils [65–68], tannin [69,70], or improving the wood properties and durability through wood chemical modification and heat treatment process [71–73] were becoming

important subjects in the wood protection field. For the liquid product from the slow pyrolysis, the potential is attractive because it encourages the valorization of the biomass waste and the liquid as the by-product. It is also worth noting that, in most cases, the organic vapors would often just be passed as waste and considered insignificant when the primary objective is producing char [74]. Further, for wood application, fixation using the minimum product through formulation should be carried out for extended efficacy through life and minimize the impacts on human health and the environment.

#### 4 Conclusion

Slow pyrolysis is an attractive method for converting the unstable biomass, such as sugarcane bagasse, which contains high volatile matter to a stable energy-rich product, even with limited yield but results in the favorable calorific value improvement. In this study, apart from its utilization for fuel due to the high carbon content (70.50%) and HHV (27.85 MJ/kg), bagasse char is also potential to be used as soil amendments, as it has a high ash recovery (up to 98%). Additionally, the recovery of liquid products results in high mass yield, which can be optimized for its potential for wood preservatives. Bagasse pyrolysis liquid contains various chemicals, such as acetic acid, formic acid, propionic acid, phenols, and guaiacol derivatives, exhibiting the anti-fungal and anti-termite properties. At low concentration (0.25%), pyrolysis liquid was able to inhibit the growth of *C. puteana* and *T. versicolor* and was effective against *R. flavipes* at a concentration of 10%. Due to its efficacy against fungi and termites, pyrolysis liquid is potentially interesting for the formulation of wood protection systems. Further study is worth carrying out, particularly in investigating the behavior of pyrolysis liquid once it is integrated into the wood given the nature of pyrolysis liquid that is acidic and has a high vapor pressure in the ambient temperature.

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**Conflicts of Interest:** The authors declare that they have no conflicts of interest to report regarding the present study.

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